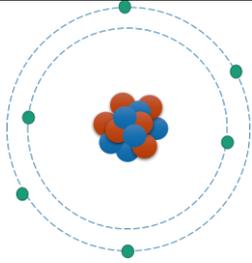
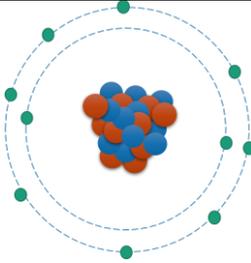
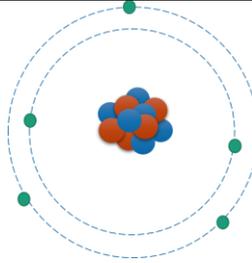
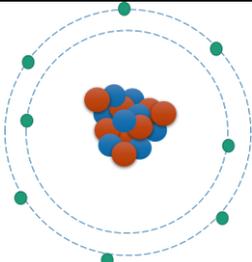
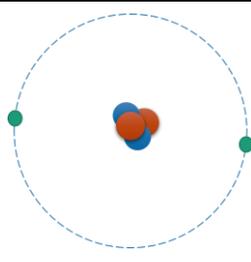
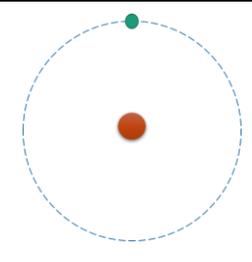


Elements and the Periodic Tables Assignment – Student Edition

I. Identifying Elements

Identify the element in the diagrams below.

1.	2.	3.
		
4.	5.	6.
		

II. Atomic Math

Fill in the chart below. Try to complete it without using the periodic table.

	Element	Atomic Number	Atomic Mass	Number of Protons	Number of Electrons	Number of Neutrons
1	Calcium	20				20
2	Tin		119	50		
3	Iodine			53		74
4	Silver		108			61
5	Neon	10	21			

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III. Ancient Elements

Some elements have kept the symbol, which was based on their Latin name, even though the element's name has been translated into English. This explains why some elements appear to have symbols that do not match their name. A description of these elements is given below.

Use this description and the periodic table to work out the element's English name and its symbol.

1. *Plumbium* is a Latin word. Historically water pipes were made from this soft, workable metal. People who had the job of working on and building water pipes were called plumbers.

Name of element _____

Symbol: _____



2. *Hydrargyrum* is Latin and means 'liquid silver'. This element is the only metal which is a liquid at room temperature. Historically it was used in thermometers.

Name of element _____

Symbol: _____



3. *Aurum* has been sort after and highly prized for millennia. It is shiny and yellow in color and used to make jewelry.

Name of element _____

Symbol: _____



4. *Ferrum* is Latin and describes one of the world's most widely used metals. It shares the name with an instrument used to press clothes.

Name of element _____

Symbol: _____



Name: _____ Period: _____ Date: _____

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5. *Kalium* is a Latin word which was derived from the Dutch word “potaschen” meaning pot ashes.

Name of element _____

Symbol: _____



IV. True or False?

Decide if the following statements are true or false. Correct the false statements to make them true.

1. The atomic number represents the number of protons and electrons combined.
2. An element contains one type of atom.
3. The number of neutrons tells us which element it is.
4. The atomic number is always larger than the atomic mass.
5. The nucleus of an atom contains protons and neutrons.

V. Interpreting the Periodic Table

Use the periodic table from your guided notes to answer the following questions.

1. Name the elements found in group 18 of the table.

2. Is helium a metal or non-metal? Give a reason for your answer.

3. How does the number of protons change as you move along a period, from left to right?

4. Name two elements from group 2 of the periodic table. _____

5. To which group does oxygen belong? _____