

Chemistry Lesson Plan

Topic	Structure and Properties of Matter
Lesson Title	Periodic Table
Lesson Number	7e
Next Generation Science Standards:	<p>HS-PS1-1. Use the periodic table as a model to predict the relative properties of elements based on the patterns of electrons in the outermost energy level of atoms.</p> <p>HS-PS1-2. Construct and revise an explanation for the outcome of a simple chemical reaction based on the outermost electron states of atoms, trends in the periodic table, and knowledge of the patterns of chemical properties.</p>
Learning objectives:	<ul style="list-style-type: none"> Use the periodic table to determine trends in electronegativity, electron affinity, atomic radius ionization energy, melting point and metallic character.
"I can" statement:	<ul style="list-style-type: none"> I can identify trends in electronegativity, electron affinity, atomic radius ionization energy, melting point and metallic character using the periodic table.
Prior Knowledge: <ul style="list-style-type: none"> Atomic and mass numbers Arrangement of the periodic table into groups. Obtaining total numbers of electrons and valence electrons 	
Vocabulary: Atomic number, atomic mass, property, groups, metal, non-metal, metalloid, valence electrons, periods, atomic radius, ionization energy, electronegativity, electron affinity, melting point, metallic character, electron shielding	
Summary of Activities: <ol style="list-style-type: none"> Distribute and complete bell ringer activity. Discuss guided notes and slideshow, with students. Complete the Vocabulary worksheet and/or graphic organizer activity (if time allows) Exit quiz 	
Additional Resources: YouTube clips on: <ul style="list-style-type: none"> Periodic trends Electronegativity 	
Homework: Homework task	
Assessment: <ul style="list-style-type: none"> Bell work Exit quiz End of unit review 	

