

Electron Configuration Homework

1. Name the elements that follow according to the following electron notation:

- a. $1s^2 2s^2 2p^4$
- b. $1s^2 2s^2 2p^6 3s^1$
- c. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- d. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- e. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- f. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
- g. $1s^2 2s^2 2p^6 3s^2 3p^4$
- h. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$

2. Give the electron configuration for the following elements:

- a. Helium
- b. Carbon
- c. Phosphorous
- d. Argon
- e. Potassium
- f. Manganese
- g. Caesium