

Chemistry Lesson Plan

Topic	Structure and Properties of Matter
Lesson Title	Electron Configuration
Lesson Number	1-6c
Next Generation Science Standards:	<p>HS-PS1-1. Use the periodic table as a model to predict the relative properties of elements based on the patterns of electrons in the outermost energy level of atoms.</p> <p>HS-PS1-2. Construct and revise an explanation for the outcome of a simple chemical reaction based on the outermost electron states of atoms, trends in the periodic table, and knowledge of the patterns of chemical properties.</p>
Learning objectives:	<ul style="list-style-type: none"> Describe how electrons occupy orbitals according to Hund's Rule Draw orbital diagrams
"I can" statement:	<ul style="list-style-type: none"> I can use Hund's Rule to construct orbital diagrams.
Prior Knowledge:	
<ul style="list-style-type: none"> Pauli Exclusion Principle Aufbau Principle S,p,d,f orbitals 	
Vocabulary:	
Atomic number, electron configuration, energy level, valence electron, Aufbau Principle, Pauli Exclusion Principle, s-orbital, p-orbital, d-orbital, f-orbital, Hund's Rule, Electron orbital diagram, sub-level,	
Summary of Activities:	
<ol style="list-style-type: none"> Distribute and complete bell ringer activity. Discuss guided notes and slideshow, with students. Complete the Vocabulary worksheet and/or doodle notes activity, Exit quiz 	
Additional Resources:	
YouTube clips on:	
<ul style="list-style-type: none"> Electron Orbitals Aufbau's Principle, Hund's Rule, Pauli Exclusion Principle 	
Homework:	
Homework task	
Assessment:	
<ul style="list-style-type: none"> Bell work Exit quiz End of unit review 	

