

Electron Configuration Bell Ringer – Student Edition

Select the best answer for each question.

- 1. Which orbitals are present in the 3rd electron shell?**
 - a. S-orbitals only.
 - b. S, p, d and f orbitals.
 - c. S and p orbitals
 - d. S, p and d orbitals.
- 2. Which of the following statements regarding orbitals is false?**
 - a. The 4d orbital has higher energy level than the 5s orbital
 - b. Orbitals can only fit a maximum of two electrons per orbital
 - c. Only the fourth energy level has d orbitals
 - d. The total number of electrons able to fit in the second orbital is 8.
- 3. The Aufbau principle states that:**
 - a. Electrons fill orbitals closer to the nucleus first as they have a lower energy levels.
 - b. Electrons always fill s-orbitals first
 - c. Electrons must fill all available energy levels
 - d. There are four different types of orbitals
- 4. Which of the following orbitals has the highest energy level?**
 - a. 4s
 - b. 4f
 - c. 4d
 - d. 4p
- 5. Which of the following has the lowest energy level?**
 - a. 5p
 - b. 6p
 - c. 4f
 - d. 5d