

Chemistry Lesson Plan

Topic	Structure and Properties of Matter
Lesson Title	Electron Configuration
Lesson Number	1- 6b
Next Generation Science Standards:	<p>HS-PS1-1. Use the periodic table as a model to predict the relative properties of elements based on the patterns of electrons in the outermost energy level of atoms.</p> <p>HS-PS1-2. Construct and revise an explanation for the outcome of a simple chemical reaction based on the outermost electron states of atoms, trends in the periodic table, and knowledge of the patterns of chemical properties.</p>
Learning objectives:	<ul style="list-style-type: none"> Name and identify the four types of orbitals as s, p, d and f Describe how electrons occupy orbitals according to the Aufbau and Pauli exclusion principles
"I can" statement:	<ul style="list-style-type: none"> I can identify the 4 types of orbitals and describe using the Aufbau and Pauli principles how they are filled.
Prior Knowledge:	
<ul style="list-style-type: none"> The structure of the atom Obtaining and using the atomic number 	
Vocabulary:	
Atomic number, electron configuration, energy level, valence electron, Aufbau Principle, Pauli Exclusion Principle, s-orbital, p-orbital, d-orbital, f-orbital,	
Summary of Activities:	
<ol style="list-style-type: none"> Distribute and complete bell ringer activity. Discuss guided notes and slideshow, with students. Complete the Vocabulary worksheet and/or doodle notes activity, Exit quiz 	
Additional Resources:	
YouTube clips on:	
<ul style="list-style-type: none"> Electron Orbitals Aufbau's Principle 	
Homework:	
Homework task	
Assessment:	
<ul style="list-style-type: none"> Bell work Exit quiz End of unit review 	

