
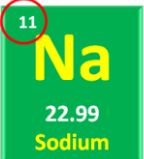


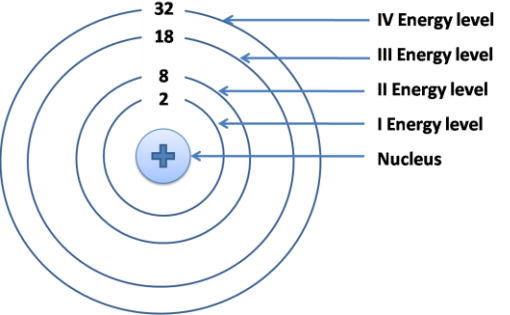
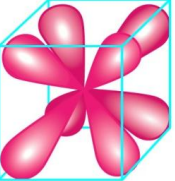




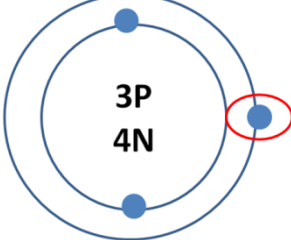


Electron Configuration Vocabulary Task Teacher Edition

Fill in the gaps in the table below

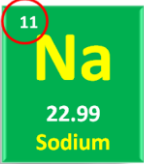


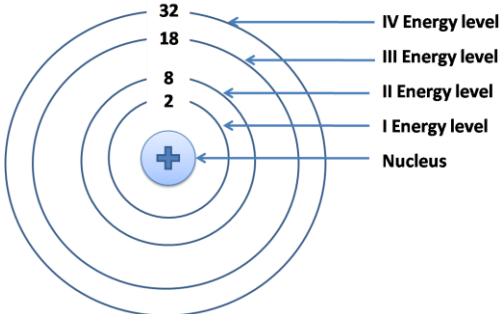
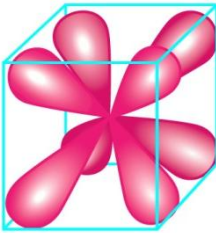

Hint/Clue	Term	Definition/Sample/Example
High  Low	Aufbau Principle	
	Atomic number	
	d-orbital	
$1s^2 2s^2 2p^3$	Electron configuration	
	Electron orbital diagram	
	Energy level	
	f-orbital	

Electron Configuration Vocabulary Task Teacher Edition




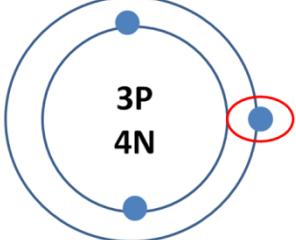
	Hund's Rule	
X^+ or X^-	Ion	
$1s^2$	Orbital notation	
	Pauli Exclusion Principle	
	p-orbital	
$[Xe]4s^1$	shorthand notation	
	s-orbital	
	Valence electron	

Electron Configuration Vocabulary Task Teacher Edition

Fill in the gaps in the table below

Hint/Clue	Term	Definition/Sample/Example
<p>High ↑ Low</p>	Aufbau Principle	The rule that states that electrons fill up the lowest energy shells before those at higher level energies.
	Atomic number	The number of protons and therefore electrons an element has.
	d-orbital	The four-lobed orbital which is found in electron shells 3-4.
$1s^2 2s^2 2p^3$	Electron configuration	The arrangement of electrons in an atom
	Electron orbital diagram	Pictorial representation of how electrons occupy the orbitals in each shell of an atom.
	Energy level	Differing distances from the nucleus of an atom which are occupied by electrons.
	f-orbital	The 8-lobed orbital found in the fourth energy level and able to hold 14 electrons
	Hund's Rule	Electrons fill up all available orbitals individually before pairing up.
X^+ or X^-	Ion	Atoms which have lost or gained electrons.

Electron Configuration Vocabulary Task Teacher Edition

$1s^2$	Orbital notation	The electron level, type of orbital and number of electrons found in a given sub-shell.
	Pauli Exclusion Principle	Electrons can only occupy the same orbital if they have the opposite spin.
	p-orbital	The 3D dumb-bell shaped orbital which occupies energy levels 2-4.
$[Xe]4s^1$	shorthand notation	Using the noble gas prior to the element to abbreviate an atoms notation
	s-orbital	Orbital found in each energy level of the atom
	Valence electron	Electrons found in the outermost energy level of an atom.