Name: _____ Period:_____ Date:_____

Bond Energy and the Types of Reactions

Bell Ringer Activity

The reaction for the production of hydrogen cyanide is as follows:

$$CH_{4(g)} + NH_{3(g)} \rightarrow HCN_{(g)} + 3H_{2(g)}$$

Bond	Bond Energy (kJ mol^{-1})
C—H	414
N—H	391
H—H	436
C≡C	860

Complete the tasks below using the bond data from the table above.

1. Draw the chemical reaction.

2. Calculate the enthalpy change for the reaction

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