
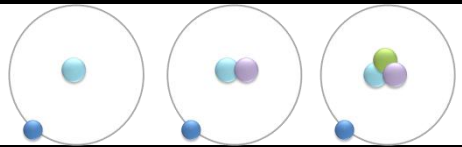

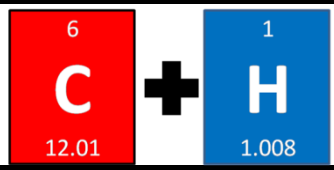



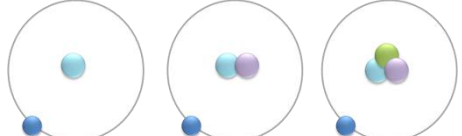

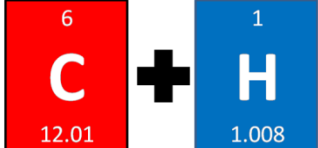
# The Mole and Molar Mass

## Vocabulary Worksheet – Teacher’s Edition

Hint	Term	Definition/Example/ Sample
<b>1 mole</b>	Avogadro’s number	
	grams	
<b>g mol<sup>-1</sup></b>	grams per mole	
	Isotope	
	mass	
	molar mass	
<b>6.02 x 10<sup>23</sup> atoms</b>	mole	
<b>?g = ?mol</b>	quantitative chemistry	
<b>A<sub>r</sub></b>	relative atomic mass	
<b>M<sub>r</sub></b>	Relative formula mass	

# The Mole and Molar Mass

## Vocabulary Worksheet – Teacher’s Edition

Hint	Term	Definition/Example/ Sample
<b>1 mole</b>	Avogadro’s number	The number ( $6.022 \times 10^{23}$ ) of atoms or molecules in a mole of any substance
	grams	The unit for mass (g)
<b>g mol<sup>-1</sup></b>	grams per mole	The unit for molar mass
	Isotope	Atoms which have the same number of protons and electrons but different numbers of neutrons
	mass	The quantity/amount of matter in a substance
	molar mass	The mass of one mole of a substance (atom, ion or compound)
<b><math>6.02 \times 10^{23}</math> atoms</b>	mole	The number of particles of any substance
<b>?g = ?mol</b>	quantitative chemistry	The study of the amounts of substances participating in a chemical reaction.
<b>A<sub>r</sub></b>	relative atomic mass	The mass of an atom relative to carbon-12
<b>M<sub>r</sub></b>	Relative formula mass	The sum of the relative atomic masses of the atoms in a compound.